

## 4.2 Chemistry Review

**A** You should be familiar with the following terms from your chemistry unit last year. Try and match each term to its definition. Cut & Paste this chart in your copybook.

Chemistry	Ion	Covalent Bond
Atom	Periodic table	Molecule
Nucleus	Particle	Compound
Neutron	Element	Diatom Molecule
Proton	Isotopes	Metal
Electron	Ionic Bond	Non-Metal
	An atom or molecule that has gained or lost one or more electrons and has a negative or positive charge.	
	A substance containing one kind of atom; the simplest form of matter.	
	A particle with a neutral charge located in the nucleus of an atom.	
	A molecule consisting of two atoms of the same element ( $H_2$ , $N_2$ , $O_2$ , $F_2$ , $Cl_2$ , $Br_2$ , $I_2$ ).	
	A chemical bond formed when two atoms share electrons in joining together.	
	The smallest part of an element that still has all the properties of that element. It contains protons, neutrons, and electrons.	
	An element that in general does not conduct electricity, is a poor conductor of heat, and is brittle when solid. Many are gases at room temperature.	
	Any of the basic units of matter and energy (molecule, atom, proton or electron).	
	Any of a group of elements that conducts heat and electricity, has a shiny luster, and is flexible.	
	The study of matter, its structures, properties, and composition, and the changes that matter undergoes	
	A chemical bond formed when one atom transfers electrons to another atom.	
	A positively charged particle in the nucleus of an atom.	
	A unit of two or more atoms held together by covalent bonds.	
	An arrangement of chemical elements based on their atomic numbers and similarity of properties.	
	A substance in which two or more elements are chemically combined.	
	The centre part of the atom.	
	Atoms of the same element having different masses due to the different numbers of neutrons in their nuclei.	
	A negatively charged particle that moves around the nucleus of an atom.	

**B** Answer the following questions in your copybook: **pg.**

**D** Copy the following figures in your copybook: **pg.**

**pg.**

**pg.**

**E** Please make sure you know the chemical symbols, names, number of sub-atomic particles and the electronic configuration for the first 20 elements in the periodic table. Cut & Paste this chart in your copybook.

Symbol	Atomic #	# p <sup>+</sup>	# e <sup>-</sup>	# n <sup>0</sup>	Name	Electronic Configuration
H	1	1	1	0	Hydrogen	1
		2				
					Lithium	
Be						
B						
					Carbon	
N						
O						
					Fluorine	
					Neon	
Na						2,8,1
Mg						
					Aluminum	
					Silicon	
P						
S						
Cl						
					Argon	
K						
					Calcium	

**\*\*\*Important\*\*\***

The knowledge and concepts in this review activity were taught in third form. This year we will build upon what is covered in this worksheet. If you are having difficulty completing these questions, it is important that you ask your teacher in class or to organize extra help.